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Add different salts to water, then watch them dissolve and achieve a dynamic

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equilibrium with solid precipitate. Compare the number of ions in solution for highly soluble NaCl to other slightly soluble salts. Relate the charges on ions to the number of ions in the formula of a salt. Calculate  $K_{sp}$  values.

**Salts & Solubility - Solubility | Salt | Solutions - PhET ...**  
Salts and Solubility 1: introduction to salts  
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Description I have  
written a series of five  
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simulation to be used  
throughout the year.  
This is the first in the  
series. I used this in  
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concentrations in  
terms of moles solute

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values ...** answers Key

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through an intuitive,  
game-like environment  
where students learn  
through exploration  
and discovery.

## **Salts and Solubility 2: Solubility (Inquiry Based) - PhET ...**

Introduction to Solubilit  
yPhETLab(created  
3/2011) We can  
express concentrations  
in terms of moles  
solute per liters of  
solution. Introduction.

When we cook pasta,

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we frequently dissolve table salt in water.

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charges on ions to the number of ions in the formula of a salt.

### **Unit 3: Solubility - Ms. Bunney's Classes**

Definition of Solubility  
Solubility is the ability of a solid, liquid, or gaseous chemical substance (referred to as the solute) to dissolve in solvent (usually a liquid) and form a solution. The solubility of a

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substance

fundamentally depends  
on the solvent used, as  
well as temperature  
and pressure.

## **Solubility | Introduction to Chemistry**

Lab CQs: 7/7/13: Salts  
and Solubility 3:  
Solution Equilibrium  
and  $K_{sp}$  (Inquiry  
Based) Trish Loeblein:  
UG-Intro HS: CQs Lab:  
7/9/13: Capacidade de  
solubilidade dos sais:

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clarice martins braga  
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7/6/13: Concentrations  
and Basic Solubility  
PhET Lab: Chris Bires:  
HS: HW Lab: 4/6/11:  
Salts & Solubility: steve  
hymes: HS: Lab: 4/7 ...

## **Salts & Solubility - Solutions, Chemical ... - PhET**

Visualizing ionic  
formulas using Salts  
and Solubility  
simulation from the  
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Learning Goals

Students will be able to:

- Determine the chemical formula by observation of ionic ratios in solutions
- Relate the simulation scale to real lab equipment through illustration and calculations

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**[PDF] Solubility Lab  
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A solution is a single-phase homogenous (evenly spread) mixture of two components called the solvent and the solute: The solvent is the chemical in larger amount - its greater number of molecules

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surround the  
dispersed,  
outnumbered solute.  
Chemists say that a  
solvent dissolves a  
solute.

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